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Chemistry (Quickstudy: Academic)

McGraw-Hill, Inc. **WORLD'S #1 IN ACADEMIC OUTLINE**

CHEMISTRY

PERIODIC TABLE OF THE ELEMENTS

ATOMIC STRUCTURE
Atomic Number, Z , is the number of protons in the nucleus. In neutral atoms, $Z = +$ the number of electrons. The atomic charge is the difference between the number of protons and the number of electrons. The atomic mass is the weighted average of the masses of the isotopes of an element.

ATOMIC QUANTUM NUMBERS AND ORBITALS
The principal quantum number, n , is the number of shells. The angular momentum quantum number, l , is the number of subshells. The magnetic quantum number, m_l , is the number of orbitals in a subshell. The spin quantum number, m_s , is the number of electrons in an orbital.

NUCLEAR CHEMISTRY
Nuclear reactions alter the nucleus. Particles emitted are alpha particles, beta particles, gamma rays, and neutrons. The mass number, A , is the sum of protons and neutrons. The atomic number, Z , is the number of protons.

MANY-ELECTRON ATOMS
Orbitals and quantum numbers describe the probability of finding an electron in an atom. The Aufbau principle, Pauli exclusion principle, and Hund's rule govern the filling of orbitals.

IONIZATION ENERGY
Ionization energy is the energy required to remove an electron from an atom. It increases across a period and decreases down a group.

ELECTRONEGATIVITY
Electronegativity is the ability of an atom to attract electrons in a chemical bond. It increases across a period and decreases down a group.

ATOMIC WEIGHTS
Atomic weights are listed at the bottom of the periodic table.

PERIODIC TABLE OF ELEMENTS
A detailed periodic table with element symbols, names, and atomic numbers.



Synopsis

Best-selling quick reference to chemistry. Chemistry essentials for students just beginning in chemistry or as an excellent reference throughout higher-level courses. 4-page laminated guide includes: • atomic structure • nuclear chemistry • atomic quantum numbers/orbitals • electronic atoms • types of matter/reactions • physical processes • nomenclature/stoichiometry • chemical interactions/bonding models • gases/thermodynamics/theories • and much more...

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Customer Reviews

These are great summaries to accompany coursework and to complement a book. By no means these are a replacement for either of those. But it's an excellent learning tool, like a "cheat sheet".

Great overview of Chemistry and the basics as well as the common topics. Well-organized and put together. And since it folds, you can easily store it in a folder, binder, or backpack. Great for a quick reference!

This helped me out a lot in my high school chemistry class. It is laid out very well and has all of the information I needed throughout the year. Chemistry wasn't easy for me but this investment definitely helped me better understand what I was doing!

Great reference. All the information I need in one place.

It has a lot of helpful formulas and it shows how to determine electron configurations. Very happy with the guide.

This was exactly what was needed for the college course. Exactly as described in the ad. Thanks so much. Me

Lots of good information, helpful for all those little things that you need to memorize. I bought it for my wife and she loves it.

This is what you need to prepare for a test, or if you are tutor like me, to help your students.

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